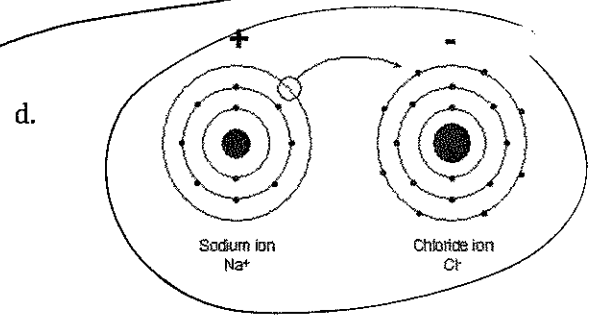
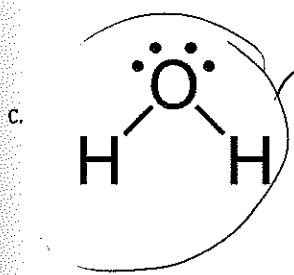
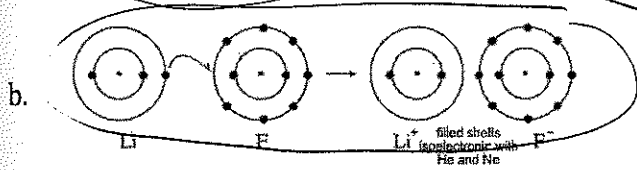


Bonding Review

Use a T Chart to divide the properties into the correct side (or on both sides):

- e⁻ shared
- e⁻ transferred
- salts that dissolve
- polar and nonpolar compounds
- single, double, or triple bonds
- formulas always simplest ratio
- cation is named 1st, anion 2nd
- bond uses valence e⁻
- uses prefixes like "di" and "tetra"
- alkanes, alkenes and alkynes
- have overall charges that = 0
- large uniform crystalline structure
- individual, independent molecules
- bonds using (whole) polyatomic ions
- bonds within polyatomic ions
- 2 nonmetals bonded
- metal and nonmetal bonded
- opposite charges attract
- uses octet rule
- often conducts electricity when dissolved
- uses Coulombic forces of attraction
- Mg(OH)₂
- CO₂
- Zn₃P₂
- NH₄⁺
- Breaks into ions when dissolved
- Does not break apart when dissolved
- Does not break apart when evaporates

Ionic	Covalent
<ul style="list-style-type: none"> e⁻ transferred salts that dissolve formulas always simplest ratio cation is 1st, anion is 2nd have overall charges = 0 large uniform (bulk) crystalline structure bonds using (whole) polyatomic ions metal + nonmetal bonded opposite charges attract often conducts electricity when dissolved breaks into ions when dissolved 	<ul style="list-style-type: none"> e⁻ shared polar and non polar compounds single, double, triple bonds uses prefixes like "di" and "tetra" alkane alkenes + alkynes individual, independent molecules bonds within polyatomic ions 2 nonmetals bonded does not break apart when dissolved does not break apart when evaporates



• Zn₃P₂

• CO₂

both

- bonds use valence electrons
- uses octet rule
- uses coulombic force of attraction
- Mg(OH)₂ (covalent between O+H)
- NH₄⁺ (covalent between N+H)